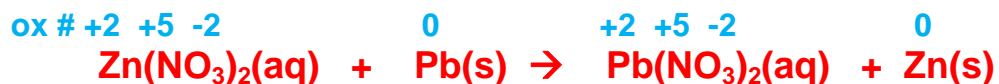


Question 1 (3 points) Give the oxidation numbers for the elements in the reaction below given that the oxidation number for oxygen is -2 and identify the reducing reagent



Pb is going from an oxidation number of 0 to +2 so it is getting more positive making it oxidized and therefore the reducing agent

Question 2 (2 points) What is the concentration 10.00 mL of a NaOH aqueous solution that is neutralized with 15 mL of a 0.4M HCl solution?

$$(\#H)(MH)(VH) = (\#OH)(MOH)(VOH)$$

$$(1)(0.4 \text{ M})(15 \text{ mL}) = (1)(MOH)(10 \text{ mL})$$

$$MOH = 0.6 \text{ M}$$