

Quiz 7 (10 points)
Dr. Taylor

Chemistry 111
Spring 2012

Name _____
First AND Last

Question 1 (5 points) How many atoms are in 34.94 grams of Zinc (MM= 65.37 g/mol)?

$$\frac{34.94 \text{ g}}{65.37 \text{ g}} \times \frac{1 \text{ mole}}{65.37 \text{ g}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.219 \times 10^{23} \text{ atoms}$$

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